

EXPERIMENTS IN NONISOTHERMAL DIFFUSION OF MOISTURE IN WOOD¹

John F. Siau

Professor of Wood Products Engineering
SUNY College of Environmental Science and Forestry, Syracuse, N.Y. 13210

Fucheng Bao

Associate Professor
Institute of Wood Industry, The Chinese Academy of Forestry
Beijing, People's Republic of China

and

Stavros Avramidis

Graduate Student
SUNY College of Environmental Science and Forestry, Syracuse, N.Y. 13210

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ABSTRACT

A series of nine experiments was performed with one surface of the wood specimen maintained at approximately 70 C and the other at 29C. The warm surface was exposed to relative humidities between 10% and 50%. One series of four experiments was conducted with the cool surfaces exposed to relative humidities between 34% and 42%, while those of the other series of five were exposed at 52% to 53%. In both series, a reversal of flux direction was observed as the relative humidity of the warm surface was increased. The results were analyzed using the general sorption data presented in the USDA Wood Handbook using two equations, the first of which is based on a gradient of activated moisture content and the second on a gradient of chemical potential that contains an additional term to account for the effect of the temperature gradient. Both equations predicted the observed reversal of flux direction, but the chemical potential equation generally provided the better fit to the experimental results.

Keywords: Nonisothermal diffusion, activated moisture content, chemical potential, activation energy of diffusion.

BACKGROUND AND THEORY

A series of four nonisothermal experiments was carried out by Siau and Babiak (1983) and these were analyzed by using the general sorption data in the Wood Handbook (USDA 1955) by applying the equation based on a gradient of activated moisture content derived by Skaar and Siau (1981). A reasonable agreement was obtained between the theoretical and experimental fluxes. A later series of five experiments was performed by Siau and Jin (1985), and these were also analyzed using the Wood Handbook data by applying both the activated-moisture-content equation and a modified one based on a gradient of chemical potential. The theoretical calculations were in fair agreement with the experimental results.

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TABLE I. Summary of data on nonisothermal diffusion experiments.

Exp. no.	B1	B2	B3	B4	C1	C2	C3	C4	C5
*T _w , °C	68.3	69.2	70.1	69.9	67.0	69.6	69.8	69.9	70.2
*T _c , °C	28.5	29.3	28.3	29.2	27.6	27.8	27.9	28.0	28.7
T, diffusion cup, °C	13.3	13.3	13.3	13.3	18.0	18.0	18.0	18.0	18.0
H _w , %	11.3	20.0	32.4	47.4	10.5	21.3	34.0	42.4	49.8
H _c , %	36.1	33.7	39.4	42.1	51.7	51.9	53.1	53.3	54.6
M _w , %	2.0	3.3	4.7	6.4	2.1	3.4	4.8	5.8	6.7
M _c , %	6.8	6.5	7.3	7.7	9.3	9.3	9.5	9.5	9.7
Exp. flux, g/cm ² s × 10 ⁹	3.7	1.4	-6.5	-13.6	12.6	8.0	4.8	3.6	-8.2

* Subscripts: w refers to warm surface of specimen and c to cool surface.

The equation based on a gradient of activated moisture content (dM^*/dx) proposed by Skaar and Siau (1981) may be written as:

$$J = -K_M \left[\frac{ME_b}{T(RT + 70M)} \frac{dT}{dx} + \frac{dM}{dx} \right] \quad (1)$$

Gradient of activated moisture content

where E_b = activation energy for bound-water diffusion, cal/mol; J = flux, g/cm² s, K_M = coefficient for bound-water diffusion based on a gradient of percent moisture content, g/(cm s %); M = percent moisture content; R = universal gas constant; T = Kelvin temperature; x = distance from cool surface, cm.

The chemical-potential equation is a modified form of Eq. (1) of Siau (1980) in which the thermal heat of transfer (Q^*) is assumed to be equal to the activation energy for bound-water diffusion (E_b). The value of this latter quantity is taken as (9,200–70 M) cal/mol in agreement with the determinations of Stamm (1959) and Choong (1963). This latter equation (Siau 1984) may be written in similar form to Eq. (1) as:

$$J = -K_M \left[\frac{H}{RT} \left(\frac{\partial M}{\partial H} \right)_T \left(\frac{d\mu_1^\circ}{dT} + \frac{E_0 + E_L + E_b}{T} + R \ln Hp_0/7600 \right) \frac{dT}{dx} + \frac{dM}{dx} \right] \quad (2)$$

Gradient of chemical potential

where E_L = molar heat of sorption of bound water, E_0 = molar heat of vaporization of free water, H = percent relative humidity, p_0 = saturated vapor pressure at T , cm Hg.

PROCEDURE

The experimental setup was similar to that described by Siau and Babiak (1983) except for significantly improved instrumentation and control of the relative humidities in contact with the wood surfaces. Temperatures were measured using a thermister readout with a resolution of 0.1 C. A fan was used to assure adequate air circulation in the measurement of the wet-bulb temperature within the controlled chamber, which was a 4½-cubic-foot Climate Lab produced by Parameter

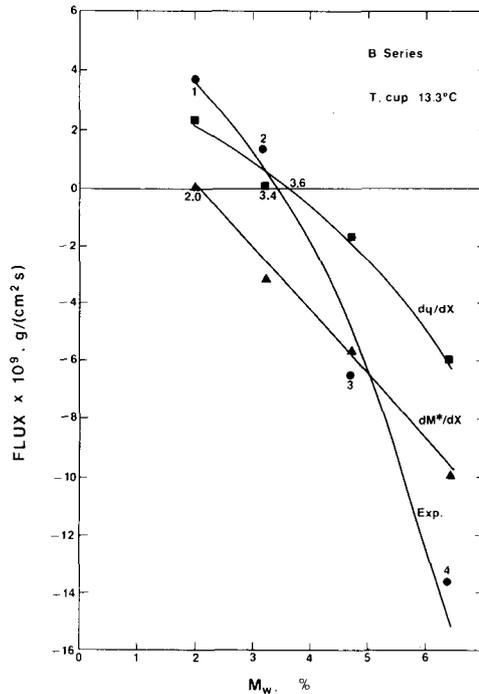


FIG. 1. A plot of experimental and theoretical fluxes vs. equilibrium moisture content on the warm side of the specimen (M_w). Theoretical results were calculated from Eqs. (1) and (2). The diffusion-cup temperature was maintained at 13.3 C. The experimental fluxes are indicated by dots that are numbered corresponding to the four experiments in the B series.

Generation and Control, Inc. The warm surface of the specimen was in equilibrium with the conditions in the chamber. The latter was placed in a room maintained at 20 ± 1 C with the relative humidity held between 30% and 50% with a dehumidifier. The relative humidity adjacent to the cool side of the test specimen was measured with lithium-chloride-coated sensors manufactured by Newport Scientific. These were previously calibrated over solutions or in the controlled chambers. The relative humidity in equilibrium with the cool surface was controlled by cooling the diffusion-cup water reservoir by means of a thermoelectric cooling chip to maintain the desired saturated vapor pressure over the water surface.

In the first series (B) this temperature was maintained at 13.3 C, while in the second series (C) it was controlled at 18.0 C. Two matched specimens of *Pinus monticola* Dougl. (western white pine) were used, with one specimen in each series. The specific gravity was 0.38 after vacuum drying to equilibrium at 50 C. The duration of the experiments was approximately two weeks.

RESULTS

The average conditions maintained during the experiments are revealed in Table 1. The diffusion-cup temperature for the B-series resulted in relative humidities

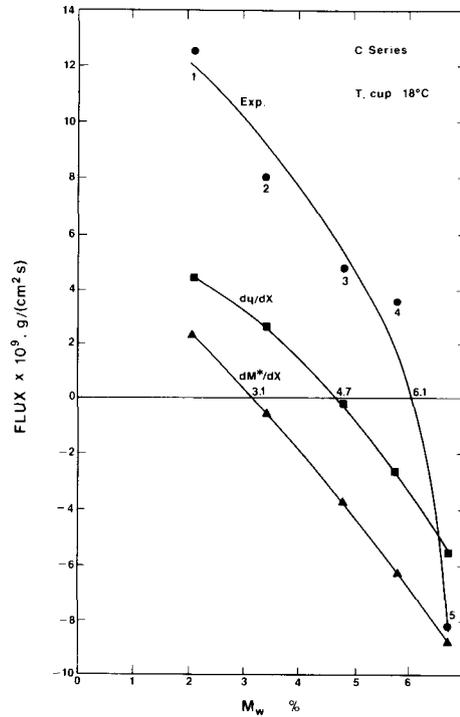


FIG. 2. Experimental and theoretical fluxes as a function of M_w for the five experiments of the C series in which the temperature of the diffusion cup was maintained at 18.0 C.

on the cool surfaces between 34% and 42%, with the variations due to differences in flux and cool-surface temperatures. The relative humidities on the warm surface were maintained between 11% and 47% for the four experiments in this series. In the five experiments of the C-series, the higher diffusion-cup temperature resulted in relative humidities between 52% and 55% on the cool surfaces. Relative humidities on the warm surfaces were varied between 11% and 50%.

In Table 2 the experimental fluxes are compared with the theoretical values calculated from Eqs. (1) and (2) by the Euler method as explained by Siau and

TABLE 2. Summary of experimental and theoretical fluxes.

Exp. no.	M_w	M_c	Flux, $\text{g}/\text{cm}^2 \text{ s} \times 10^9$		
			Exp.	dM^*/dx	$d\mu/dx$
B1	2.0	6.8	3.7	0.0	2.4
B2	3.3	6.5	1.4	-3.1	0.1
B3	4.7	7.3	-6.5	-5.7	-1.7
B4	6.4	7.7	-13.6	-10.0	-6.0
C1	2.1	9.3	12.6	2.4	4.6
C2	3.4	9.3	8.0	-0.5	2.7
C3	4.8	9.5	4.8	-3.7	-0.1
C4	5.8	9.5	3.6	-6.3	-2.6
C5	6.7	9.7	-8.2	-8.7	-5.5

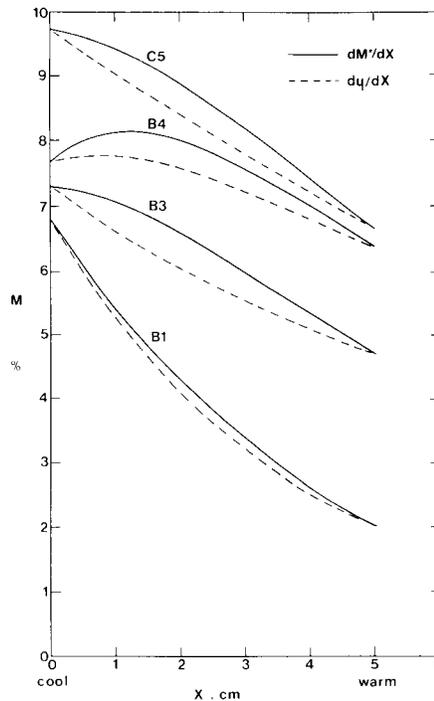


FIG. 3. Moisture profiles in Experiments B1, B3, B4, and C5 as calculated from Eqs. (1) and (2).

Babiak (1983). These results are presented graphically in Figs. 1 and 2, in which the flux is plotted as a function of the moisture content on the warm surface (M_w) for each series. The plots were made in this manner because the moisture contents on the cool side (M_c) were maintained approximately constant in each series.

Figure 1 indicates that the experimental flux changed direction at a value of M_w of 3.4%. Calculations based on a gradient of activated moisture content result in expected flux reversal of $M_w = 2.0\%$, while those based on chemical potential predicted reversal at $M_w = 3.6\%$. The flux values calculated from Eq. (2) were in better agreement with the experiments at lower values of M_w while Eq. (1) gave better agreement at higher values of M_w .

Figure 2 reveals that, in the C series, experimental flux reversal occurred at $M_w = 6.1\%$, while Eq. (1) based on activated moisture content predicted reversal at $M_w = 3.1\%$ and Eq. (2) gave a value of 4.7%. The values of flux calculated from Eq. (2) were also closer to experiment than those from Eq. (1) except for Experiment C-5.

It is concluded that Eq. (2) based on chemical potential gave a better agreement with the experimental results than Eq. (1) in both the B and C series. It must be realized, however, that this conclusion is based on the use of the general Wood Handbook sorption data, which may not be typical of the western white pine specimens used in these experiments. It would be much more meaningful to base the comparison on our own experimental sorption data from specimens closely matched to those used in the experiments. Such sorption data are now being accumulated and their use should lend much more validity to future results.

Plots of the moisture profiles calculated from Eqs. (1) and (2) are revealed in Fig. 3. In all cases the activated-moisture equation gave higher values of M for a given value of x than did the chemical-potential equation in agreement with the results of Siau and Jin (1985). It is interesting to note that both equations predicted a maximum value of M within the specimen in Experiment B4. There was also a large increase in experimental flux between B3 and B4 corresponding to a relatively small increase in H_w from 32.4% to 47.4%. When H_w was increased to 63% in Exp. B5, a much greater flux of -46×10^{-9} g/cm² s occurred, which caused a loss of control of the relative humidity on the cool side of the specimen and condensation in the plastic tube adjacent to the diffusion cup, which made accurate monitoring of the experiment impossible. An excessive time (over one month) was required to attain a constant flux. In addition to this, a much higher theoretical peak moisture content within the specimen was calculated than in B4, leading to the conclusion that condensation had probably occurred within the specimen. It was therefore concluded that there is a practical upper limit to the relative humidity in equilibrium with the warm surface of the specimen in a nonisothermal diffusion experiment. In the B-series, this upper limit was probably between 45% and 60%. For these reasons, the results of Experiment B5 were omitted from this analysis.

Further similar experimentation is planned with a greater range of partial vapor pressures on the cool side, along with sorption studies of specimens matched to those used in the nonisothermal experiments.

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